

# The Modern Periodic Table of the Elements

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Hydrogen 1 <b>H</b> 1.01 2.1	2											Helium 2 <b>He</b> 4.00 ---						
Lithium 3 <b>Li</b> 6.94 1.0	Beryllium 4 <b>Be</b> 9.01 1.5											Boron 5 <b>B</b> 10.81 2.0	Carbon 6 <b>C</b> 12.01 2.5	Nitrogen 7 <b>N</b> 14.01 3.0	Oxygen 8 <b>O</b> 16.00 3.5	Fluorine 9 <b>F</b> 19.00 4.0	Neon 10 <b>Ne</b> 20.18 ---	
Sodium 11 <b>Na</b> 22.99 0.9	Magnesium 12 <b>Mg</b> 24.31 1.2											Aluminum 13 <b>Al</b> 26.98 1.5	Silicon 14 <b>Si</b> 28.09 1.8	Phosphorus 15 <b>P</b> 30.97 2.1	Sulfur 16 <b>S</b> 32.07 2.5	Chlorine 17 <b>Cl</b> 35.45 3.0	Argon 18 <b>Ar</b> 39.95 ---	
Potassium 19 <b>K</b> 39.10 0.8	Calcium 20 <b>Ca</b> 40.08 1.0	3	4	5	6	7	8	9	10	11	12	Gallium 31 <b>Ga</b> 69.72 1.6	Germanium 32 <b>Ge</b> 72.61 1.8	Arsenic 33 <b>As</b> 74.92 2.0	Selenium 34 <b>Se</b> 78.96 2.4	Bromine 35 <b>Br</b> 79.90 2.8	Krypton 36 <b>Kr</b> 83.80 3.0	
Rubidium 37 <b>Rb</b> 85.47 0.8	Strontium 38 <b>Sr</b> 87.62 1.0	Yttrium 39 <b>Y</b> 88.91 1.2	Zirconium 40 <b>Zr</b> 91.22 1.4	Niobium 41 <b>Nb</b> 92.91 1.6	Molybdenum 42 <b>Mo</b> 95.94 1.8	Technetium 43 <b>Tc</b> (98) 1.9	Ruthenium 44 <b>Ru</b> 101.07 2.2	Rhodium 45 <b>Rh</b> 102.91 2.2	Palladium 46 <b>Pd</b> 106.42 2.2	Silver 47 <b>Ag</b> 107.87 1.9	Cadmium 48 <b>Cd</b> 112.41 1.7	Indium 49 <b>In</b> 114.82 1.7	Tin 50 <b>Sn</b> 118.71 1.8	Antimony 51 <b>Sb</b> 121.76 1.9	Tellurium 52 <b>Te</b> 127.60 2.1	Iodine 53 <b>I</b> 126.90 2.5	Xenon 54 <b>Xe</b> 131.29 2.6	
Cesium 55 <b>Cs</b> 132.91 0.7	Barium 56 <b>Ba</b> 137.33 0.9	57-70 *	Lutetium 71 <b>Lu</b> 174.97 1.1	Hafnium 72 <b>Hf</b> 178.49 1.3	Tantalum 73 <b>Ta</b> 180.95 1.5	Tungsten 74 <b>W</b> 183.84 1.7	Rhenium 75 <b>Re</b> 186.21 1.9	Osmium 76 <b>Os</b> 190.23 2.2	Iridium 77 <b>Ir</b> 192.22 2.2	Platinum 78 <b>Pt</b> 195.08 2.2	Gold 79 <b>Au</b> 196.97 2.4	Mercury 80 <b>Hg</b> 200.59 1.9	Thallium 81 <b>Tl</b> 204.38 1.8	Lead 82 <b>Pb</b> 207.20 1.8	Bismuth 83 <b>Bi</b> 208.98 1.9	Polonium 84 <b>Po</b> (209) 2.0	Astatine 85 <b>At</b> (210) 2.2	Radon 86 <b>Rn</b> (222) 2.4
Francium 87 <b>Fr</b> (223) 0.7	Radium 88 <b>Ra</b> (226) 0.9	89-102 **	Lanthanum 103 <b>Lr</b> (262) ---	Rutherfordium 104 <b>Rf</b> (261) ---	Dubnium 105 <b>Db</b> (262) ---	Seaborgium 106 <b>Sg</b> (263) ---	Bohrium 107 <b>Bh</b> (262) ---	Hassium 108 <b>Hs</b> (265) ---	Meitnerium 109 <b>Mt</b> (266) ---	Ununnilium 110 <b>Uun</b> (271) ---	Unununium 111 <b>Uuu</b> (272) ---	Ununbium 112 <b>Uub</b> (277) ---	Ununtrium 113 <b>Uut</b> (284) ---	Ununquadium 114 <b>Uuq</b> (289) ---	Ununpentium 115 <b>Uup</b> (288) ---	Ununhexium 116 <b>Uuh</b> (291) ---	Ununoctium 118 <b>Uuo</b> (294) ---	

Average relative masses are 2001 values, rounded to two decimal places.

All average masses are to be treated as measured quantities, and subject to significant figure rules. Do not round them further when performing calculations.

Element name → Mercury

Atomic # → 80

Symbol → Hg

Avg. Mass → 200.59

Electronegativity → 1.9

\*lanthanides

Lanthanum 57 <b>La</b> 138.91 1.1	Cerium 58 <b>Ce</b> 140.12 1.1	Praseodymium 59 <b>Pr</b> 140.91 1.1	Neodymium 60 <b>Nd</b> 144.24 1.1	Promethium 61 <b>Pm</b> (145) 1.1	Samarium 62 <b>Sm</b> 150.36 1.2	Europium 63 <b>Eu</b> 151.97 1.1	Gadolinium 64 <b>Gd</b> 157.25 1.2	Terbium 65 <b>Tb</b> 158.93 1.1	Dysprosium 66 <b>Dy</b> 162.50 1.2	Holmium 67 <b>Ho</b> 164.93 1.2	Erbium 68 <b>Er</b> 167.26 1.2	Thulium 69 <b>Tm</b> 168.93 1.3	Ytterbium 70 <b>Yb</b> 173.04 1.1
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\*\*actinides

Actinium 89 <b>Ac</b> (227) 1.1	Thorium 90 <b>Th</b> 232.04 1.3	Protactinium 91 <b>Pa</b> 231.04 1.5	Uranium 92 <b>U</b> 238.03 1.4	Neptunium 93 <b>Np</b> (237) 1.4	Plutonium 94 <b>Pu</b> (244) 1.3	Americium 95 <b>Am</b> (243) 1.3	Curium 96 <b>Cm</b> (247) 1.3	Berkelium 97 <b>Bk</b> (247) 1.3	Californium 98 <b>Cf</b> (251) 1.3	Einsteinium 99 <b>Es</b> (252) 1.3	Fermium 100 <b>Fm</b> (257) 1.3	Mendelevium 101 <b>Md</b> (258) 1.3	Nobelium 102 <b>No</b> (259) 1.3
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<p><b><u>Atomic structure and energy</u></b></p> $\Delta E = h\nu$ $c = \lambda\nu$ <p><b><u>Gases, liquids and solutions</u></b></p> $PV = nRT$ $n = \frac{m}{M}$ $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ $P_{TOTAL} = P_A + P_B + P_C + \dots$ $K = ^\circ C + 273$ $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$ <p>Kinetic Energy (KE) = <math>\frac{1}{2}mv^2</math></p> $\text{Density} = \frac{m}{V}$ $\Delta T_f = iK_f \cdot \text{molality}$ $\Delta T_b = iK_b \cdot \text{molality}$ <p><b><u>Acids, Bases, and pH</u></b></p> $K_w = [H^+][OH^-] = 1.0 \times 10^{-14} \text{ (at } 25^\circ C)$ $pH = -\log [H^+] \quad pOH = -\log [OH^-]$ $[H^+] = 10^{-pH} \quad [OH^-] = 10^{-pOH}$	<p><b><u>Equilibrium</u></b></p> $Q = \frac{[C]^c [D]^d}{[A]^a [B]^b} \text{ where } aA + bB \rightleftharpoons cC + dD$ <p><b><u>Thermochemistry</u></b></p> $\Delta H^0 = \sum \Delta H_f^0 \text{ products} - \sum \Delta H_f^0 \text{ reactants}$ $\Delta G^0 = \Delta H^0 - T\Delta S^0$ $q = mc\Delta T \quad C_p = \frac{\Delta H}{\Delta T}$ <p><b><u>Constants</u></b></p> <p>Speed of light, <math>c = 3.00 \times 10^8</math> meters/s</p> <p>Planck's Constant, <math>h = 6.63 \times 10^{-34}</math> joule·s</p> <p>Avogadro's Number = <math>6.022 \times 10^{23}</math></p> <p>Gas Constant, <math>R = 0.0821 \frac{L \cdot atm}{mol \cdot K}</math></p> <p>STP = <math>0.000^\circ C</math> and <math>1.000</math> atmosphere</p> <p>Standard molar volume = <math>22.4</math> L</p> <p>Freezing point depression constant for water,</p> $K_f = \frac{1.86^\circ C}{molal}$ <p>Boiling point elevation constant for water,</p> $K_b = \frac{0.51^\circ C}{molal}$	<p><b><u>Symbols</u></b></p> <p><math>E</math> = energy</p> <p><math>\lambda</math> = wavelength</p> <p><math>\nu</math> = frequency</p> <p><math>m</math> = mass</p> <p><math>M</math> = molar mass in grams per mole</p> <p><math>q</math> = heat</p> <p><math>P</math> = pressure</p> <p><math>V</math> = volume</p> <p><math>n</math> = moles</p> <p><math>T</math> = temperature</p> <p><math>D</math> = density</p> <p><math>v</math> = velocity</p> <p><math>r</math> = rate of effusion</p> <p><math>t</math> = time (seconds = s)</p> <p><math>c</math> = specific heat capacity</p> <p><math>C_p</math> = molar heat capacity at constant <math>P</math></p> <p><math>i</math> = van't Hoff factor</p> <p><math>Q</math> = reaction quotient</p> <p><math>S^0</math> = standard entropy</p> <p><math>H^0</math> = standard enthalpy</p> <p><math>G^0</math> = standard free energy</p>
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