

- 1) What idea is the KMT based on?
- 2) What is an ideal gas?
- 3) What makes a real gas different than an ideal gas?
- 4) Describe the conditions under which a real gas is most likely to behave ideally.
- 5) If real gases do not behave ideally, what is the point of using something like the "Ideal Gas Law?"
- 6) State the five basic assumptions of the KMT. Put them in your own words!

- 7) How does the KMT explain the pressure exerted by gases?
- 8) How do gases compare with liquids and solids in terms of the distance between their molecules?
- 9) What are the five main properties of gases?
- 10) What happens to gas particles when a gas is compressed?
- 11) What happens to gas particles when a gas is heated?
- 12) What is the relationship between the temperature, speed, and kinetic energy of gas molecules?
- 13) Which of the following conditions would give you the greatest kinetic energy? Why?
a. 0°C and 2 atm b. 100 °C and 1 atm c. 75°C and 4 atm