

- Gas particles zoom around \_\_\_\_\_
  - Helium moves at \_\_\_\_\_ miles/hour
  - Normal air moves at \_\_\_\_\_ miles /hour
- Pressure comes from the gas particles \_\_\_\_\_
- Diagram of what causes pressure:

- Units of pressure

Unit	Symbol	Definition/Relationship
	Pa	SI Unit of pressure
Millimeter of mercury		How much pressure it takes to hold up a 1 mm column of mercury in a barometer.
	torr	1 torr =
Atmosphere		Average atmospheric pressure at sea level and 0°C $1 \text{ atm} = \text{_____ mmHg}$ $= \text{_____ torr}$ $= 1.01325 \times 10^5 \text{ Pa}$ $= 101.325 \text{ kPa}$ $= \text{_____ psi}$
Pound per square inch		$1 \text{ psi} = 6.89286 \times 10^3 \text{ Pa}$ $1 \text{ atm} = 14.700 \text{ psi}$

- Converting Units Example Problem #1 - How many atmospheres is 57.4 psi ?
  
- Converting Units Example Problem #2 - How many mm Hg is 245.8 psi?