

Kinetic Molecular Theory

- Kinetic Molecular Theory is abbreviated as _____
 - It is based on the idea that _____

The Kinetic Molecular Theory applies to what type of gases?

- _____
 - A hypothetical gas that follows _____
 - Doesn't really exist!!!
 - Allows us to estimate a lot of things, but they are not exactly real
- _____ can behave like ideal gases, but only under certain conditions
 - _____
 - _____

5 assumptions of KMT

- 1) Gases consist of large numbers of tiny particles that are far _____

 - Most of the volume of a gas is _____.
 - Gas particles themselves, are so small they don't actually have _____
 - Gas particles are further apart than in _____
- 2) Collisions between gas particles and between particles and container walls are _____

 - When two molecules collide with each other, they _____ their kinetic energy, but they don't _____ any energy overall.
- 3) Gas particles are in continuous, rapid, _____. They therefore, possess kinetic energy, which is energy of motion.
- 4) There are no _____ between gas particles.
 - They act like _____

5) The _____ of a gas depends on the average kinetic energy of the particles of the gas.

- _____ Temperature = _____ velocity = _____ kinetic energy

Properties of Gases

1) Expansion

- No _____
- They expand to _____ they are in

2) Fluidity

- There are no _____ between particles, so they glide/flow past each other

3) Low Density

- They are _____ so they have a low density

4) Compressibility

- Because the molecules are far apart, they can be compressed/squished

5) Diffusion

- Gas particles fly around and _____